

Chemistry Revision Test

Time : 45 mins

Max Marks: 30

General Instructions:

1. Question numbers 1 to 5 are one-mark questions. These are to be answered in one word or in one sentence.
2. Question numbers 6 to 10 are two-mark questions. These are to be answered in about 30 words each.
3. Question numbers 10 to 15 are three-mark questions. These are to be answered in about 50 words each.

Q1. The electron releasing tendency of zinc is _____ than that of copper

- (a) More
- (b) Less
- (c) equal
- (d) None of the above

Q2. What is electrolytic dissociation.

Q3. Why cane sugar doesn't conduct electricity?

Q4. Why H^+ ions are discharged more easily than Na^+ ions during electrolysis.

Q5 Why copper vessel is not used to store Silver nitrate solution.

Q6 Write the reaction at anode during the electrolysis of fused lead bromide.

Q7. How is exothermic reaction different from an endothermic reaction?

Q8. Concentrated HCL in pure liquid state does not conduct electricity. Give reason.

Q9. What are strong and weak electrolytes. Give two examples of each?

Q10. Name the electrode at the anode, cathode and electrolyte used in the electroplating for the silver spoon.

Q11. Why Aluminum is extracted by electrolytic reduction..

Q12. Ammonia is not electrolyte but ammonium hydroxide is weak electrolyte. Give reason.

Q13. CCl_4 is liquid state but doesn't conduct electricity. Why is it so?

Q14. Explain the dissociation of sodium chloride during electrolysis.

Q15. Give two differences between ionization and dissociation.