

Chemistry Revision Test 1

Time : 45 mins

Max Marks: 30

General Instructions:

1. Question numbers 1 to 5 are one-mark questions. These are to be answered in one word or in one sentence.
2. Question numbers 6 to 10 are two-mark questions. These are to be answered in about 30 words each.
3. Question numbers 10 to 15 are three-mark questions. These are to be answered in about 50 words each.

Q1. The chemical formula of lead sulphate is

- (a) Pb_2SO_4
- (b) $\text{Pb}(\text{SO}_4)_2$
- (c) PbSO_4
- (d) $\text{Pb}_2(\text{SO}_4)_3$

Q2. Which information is not conveyed by a balanced chemical equation?

- (a) Physical states of reactants and products
- (b) Symbols and formulae of all the substances involved in a particular reaction
- (c) Number of atoms/molecules of the reactants and products formed
- (d) Whether a particular reaction is actually feasible or not

Q3. Chemically rust is

- (a) hydrated ferrous oxide
- (b) only ferric oxide
- (c) hydrated ferric oxide
- (d) none of these

Q4. Both CO_2 and H_2 gases are

- (a) heavier than air
- (b) colourless
- (c) acidic in nature
- (d) soluble in water

Q5 Which of the following gases can be used for storage of fresh sampel of an oil for a long time?

- (a) Carbon dioxide or oxygen
- (b) Nitrogen or helium
- (c) Helium or oxygen
- (d) Nitrogen or oxygen

Q6 Explain the process of corrosion and rusting.

3 marks

Q7. How is exothermic reaction different from an endothermic reaction?

Q8. Give two uses of quick lime.

3 marks

Q9. Why should a magnesium ribbon be cleaned before it is burnt in air?

Q10. Write a balanced chemical equation with state symbols for the following reactions.

- (a) Solutions of barium chloride and sodium sulphate in water react to give insoluble barium sulphate and the solution of sodium chloride.
- (b) Sodium hydroxide solution (in water) reacts with hydrochloric acid solution (in water) to produce sodium chloride solution and water.

Q11. Why are decomposition reactions called the opposite of combination reactions? Write equations for these reactions.

Q12. Write one equation each for decomposition reactions where energy is supplied in the form of heat, light or electricity.

Q13. Explain the following in terms of gain or loss of oxygen with two examples each.

- (a) Oxidation
- (b) Reduction

Q14. Oil and fat containing food items are flushed with nitrogen. Why?

Q15. What are the properties and uses of bleaching powder?